



:

: 6750

: 6496

: .

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	1
	:	2
1.	3
1.1.	3
1.1.1.	3
1.1.2.	4
1.1.3.	μ	4
1.1.4.	μ	7
1.2.	9
1.3.	14
1.4.	18
1.4.1.	()	21
2.	27
2.1.	27
2.2.	27
2.2.1.	28
2.2.2.	29
2.2.2.1.	30
2.3.	31
2.3.1.	μ	31
2.4.	32
3.	34
3.1.	34
3.1.	34
3.1.1.	34
3.1.2.	35
3.1.3.	37
3.1.1.	37
3.2.	38
3.2.1.	$\mu \mu \mu \mu$	39
3.2.2.	41
3.2.3.	μ	42

. . .

3.2.4.	μ	43	
3.2.5.		44	
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3.2.7.		45	
3.3.		46	
3.3.1.	μ	μ	46
3.3.2.		47	
3.3.3.		50	
3.4.		51	
3.5.		51	
:	μ	53	
4.		54	
4.1.		54	
4.2.		55	
5.	–	59	
		62	
		63	
		69	
OLIVE MILL WASTEWATERS TOTAL ORGANIC CARBON DEGRADATION USING TiO ₂ NANOPARTICLES			69	
		79	
		()	80	
		80	

. . .

1: To 2800 . . . ,	(μ).....	3
2:		μ	4
3:			8
4:			8
5:			19
6:		μ	19
7:			21
8:			μ	28
9:			29
10: μ		μ	30
.....				30
11: μ		μ	55
12: μ		μ	55
13: 0,10 gr			TiO ₂	56
14: 0,50 gr			TiO ₂	56
15: 0,80 gr			TiO ₂	56
16:		μ	μ	56
17: μ		μ	57

. . .

μ 1:	9
μ 2:	10
μ 3:	μ	10
μ 4:	20	11
μ 5:	μ μ	12
μ 6:	μ	12
μ 7:	μμ) μμ))	13
μ 8:	15
μ 9:	16
μ 10:	33
μ 11:	μ	35
μ 12:	μ	36
μ 13:	μ	37
μ 14:	μ μ () () μ ..	38
μ 15:	μ μ μ μ ..	39
μ 16:	42
μ 17:	μ μ	43
μ 18:	μ	44
μ 19:	45
μ 20:	μ μ	46
μ 21:	μμ μ μ μ	49
μ 22:	50
μ 23:	μ μ TiO ₂	59
μ 24:	μ TOC μ	60
μ 25:	μ μ μ μ 300	61
μ 26:	μ μ μ μ / μ , μ	62

. . .

1:	(1.000)	5
2:	μ (2004)	6
3:		16
4:		17
5:		20
6:		22
7:	μ ()	24
8:	()	24
9:	μ μ μ	40
10:	μ	41
11:		52
12:	μ μ	57

...

ΜΕΡΟΣ Α'

...



3:
[: <https://olivestories.eu/>]

_____:

, μ μ (1985) (μ , 2011):

➤

•

μ

➤

•

μ

μ

(0,5 μ)

μ

➤

•

μ μ

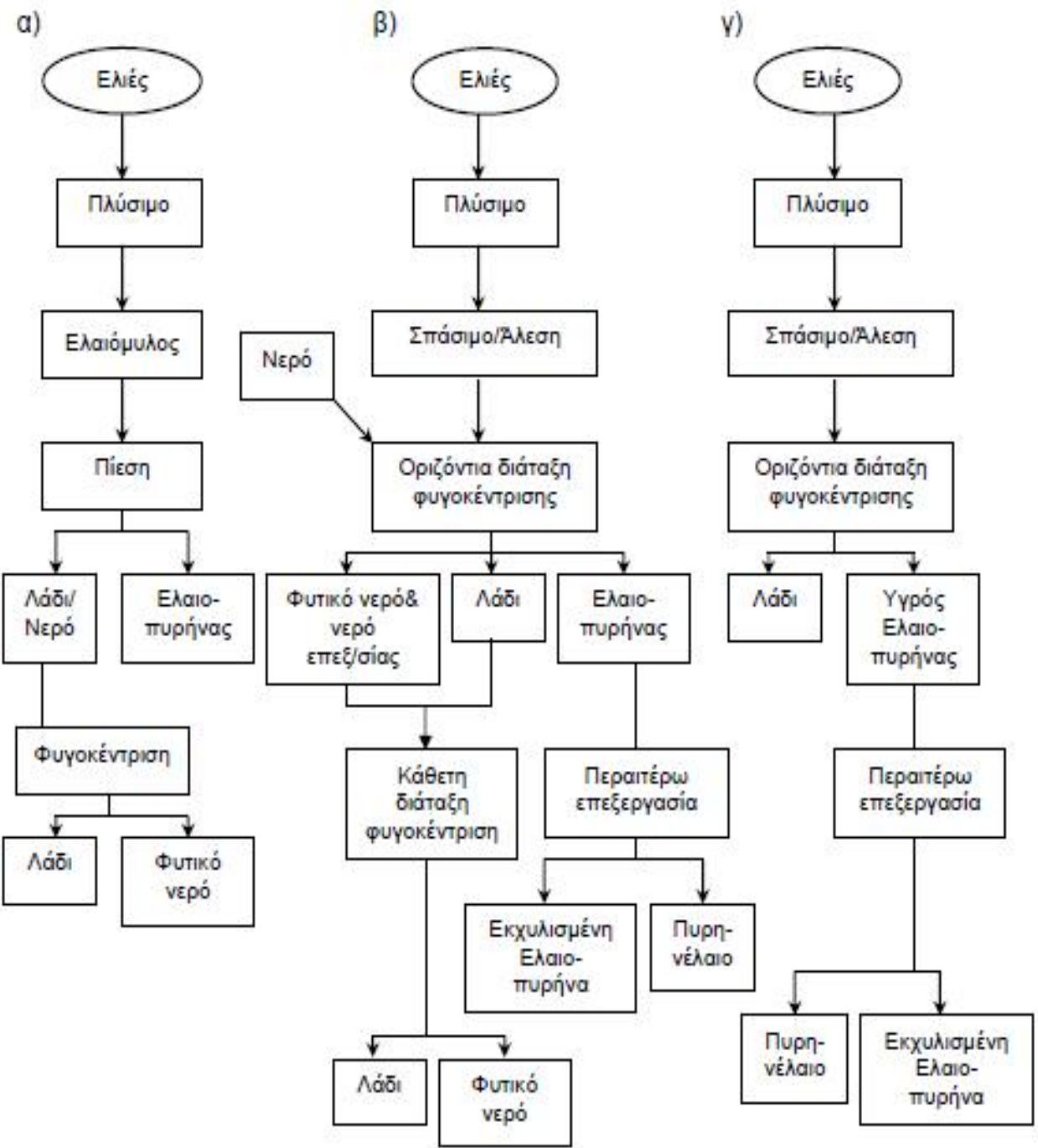
μ

1,5 μ

μ



4:
[: <https://olivestories.eu/>]



μ 7: μμ) [:) , 2009]) .

...

μ μ μ μ μ μ μ μ



5:
[: <http://www.prosodol.gr>]



6: μ
[: <http://www.prosodol.gr>]

(μ , 2011):

-
-
-
-

μ

...

-
-
-

μ

μ

(5).

5:

[: 2006]

	μ	μ	3-	2-
%	27,2 ± 1,048	50,23 ± 1,935	56,80 ± 2,188	
%	8,72 ± 3,254	3,89 ± 1,449	4,65 ± 1,736	
%	4,77 ± 0,024	3,43 ± 0,017	2,87 ± 0,014	
O %	1,38 ± 0,016	0,99 ± 0,012	0,83 ± 0,010	
μ %	24,1 ± 0,283	17,37 ± 0,203	14,54 ± 0,170	
%	11,0 ± 0,608	7,92 ± 0,438	6,63 ± 0,366	
%	2,36 ± 0,145	1,70 ± 0,105	1,42 ± 0,088	
, %	14,1 ± 0,291	10,21 ± 0,209	8,54 ± 0,175	
Kjendahl, %	0,71 ± 0,010	0,51 ± 0,007	0,43 ± 0,006	
P ₂ O ₅ , %	0,07 ± 0,005	0,05 ± 0,004	0,04 ± 0,003	
, %	1,14 ± 0,06	0,326 ± 0,035	2,43 ± 0,15	
K ₂ O, %	0,54 ± 0,045	0,39 ± 0,033	0,32 ± 0,027	
CaO, %	0,61 ± 0,059	0,44 ± 0,043	0,37 ± 0,036	
, %	42,9 ± 3,424	29,03 ± 2,317	25,37 ± 2,025	
C/N	60,7 ± 5,352	57,17 ± 5,033	59,68 ± 5,254	
C/P	588,7 ± 51,25	552,9 ± 48,20	577,2 ± 50,31	

μ

μ

μ

μ μ

μ

μ

μ

μ

μ

•

μ

μ

μ

μ

•

μ

μ

μ

μ

μ

μ

μ

μ

μ

μ

μ

μ

μ

,

μ

μ

μ

μ

μ

6: [: , 2009]

μ	acqua di vegetazione/
	água ruça/
	alpechín/
	amorca/
	amorgi (μ)/
	amurca/
	mourga (μ)/
	kara su/
	katsigaros()/
	liozumia (μ)/
	marginè/
	morga/
	mrar/
	murga/
	olive lees/
	veget abilna voda/
	vegetable water/
	olive mill wastewater/
μ	acqua di vegetazione/
	água ruça de 3 fases/
	alpechín /
	amorca/
	kara su/
	katsigaros()/
	liozumia (μ)/
	marginè/
	morga/
	mrar/
	murga/
veget abilna voda/	
μ	zubar/
	água residual de 2 fases/
	alpechín -2/
	marginè-2/
μ	jamila-2/
	mrar/

...

μ

7: [: Ursinos and Padilla, 1992; Hamdi and Ellouz, 1992] ()

μ		μ
%		83-94
%	%	4-16
%		1-2
(g/cm ³)		1,024
μ (μS/cm)		80.000-160.000
pH		4,5-6,5
μ	μ (BOD ₅) mg/L	14.000-110.0000
μ	μ (COD) mg/L	41.400-130.000

8: [: Zervakis and Balis, 1996] ()

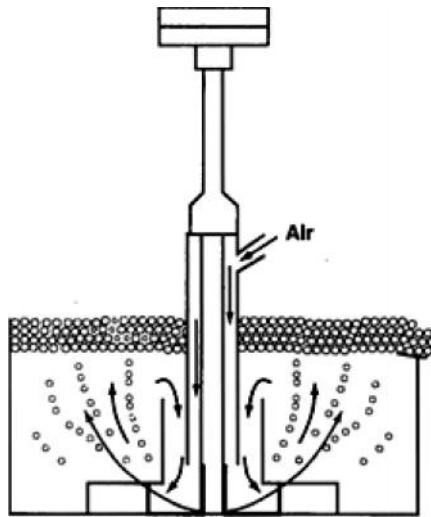
μ		(%)
		83-92
	0,03-1,00	μμ
	1,2-2,4	μ , , , ,
		, , , , ,
		, μ . .
	2,0-8,0	, , , μ ,
		, , , , ,
	0,5-1,5	, , , , ,
		μ , , , , ,
	0,5-1,5	, , , , ,
	0,4-1,5	, , , , ,
		: , , , , ,
		: , , , μ , ,
	0,3-0,8	, - , μ , ,
		, , , , ,
		, , , , ,
	0,4-1,5	K, P, Na, Ca, Mg, Fe, Mn, Zn, Cu, Cl, S

. . .

μ μ Pseudomonas μ μ .
μ μ Saccharomyces μ Penicillium ,
μ μ
Aspegillus. μ , Aerobacter, Escherichia, Bacillus, Rhizopus, Alternaria,
Fusarium (, 2009).

...

- (, 2009):
- (foaming chemicals):
- (,)
- :
- (collectors).
- : pH,
- ...



12: [: <https://www.911metallurgist.com/>]

(dissolved air flotation),

() , /

(, 2009 ; , 2010).

Actinobacteria,

(Perez et al., 2001).

3.1.3.

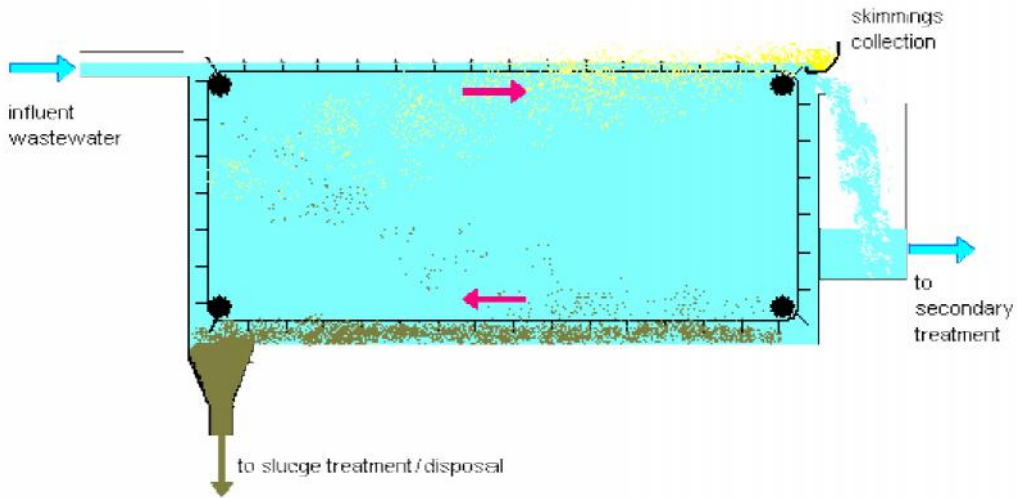
(sedimentation)

.

35% BOD₅. 50%

. (13).

. (, 2009).



μ 13: μ
[: <http://isu.indstate.edu/ebermudez/hlth210/lessonfourc.html>]

3.1.1.

μ (degreasing).

...

(Brenes et al., 1993).
 (2009)
 (nanofiltration).
 1000 mol/gr.
 10: [: , 2009]

			(%)
(No/mL)	108	0	100
(mg/L)	1090	0	100
COD (mg/L)	8950	705	92
BOD5 (mg/L)	5970	500	92
(mg/L)	150	0	100
(mg/L)	150	0	100

« μ » (flash evaporator),
 75%.
 (Dangel et al., 1995).

3.2.2.

(incineration) μ μ
 μ μ
 μ μ μ μ

...

(static incineration chambers) (fluid bed oven) (rotary tube ovens)

2009).

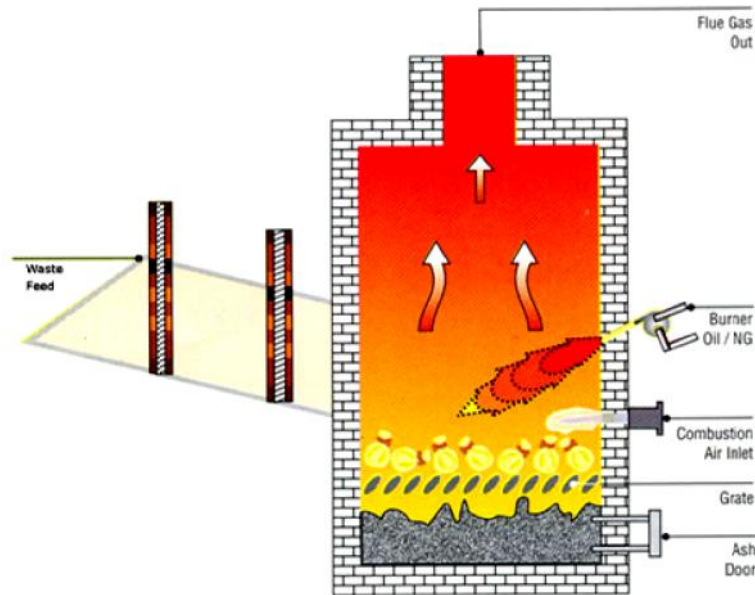


Figure 16: [Source: <http://www.maharashtradirectory.com/>]

The temperature range is 650°C to 1600°C. The feed rate is 10,000 SSU (Spermin, 2011).

(Rozzi et al., 1996; Brenes et al., 1993).

3.2.3.

(evaporation and distillation)

()

. . .

(Rozzi et al., 1996).

(, 2009). 50%

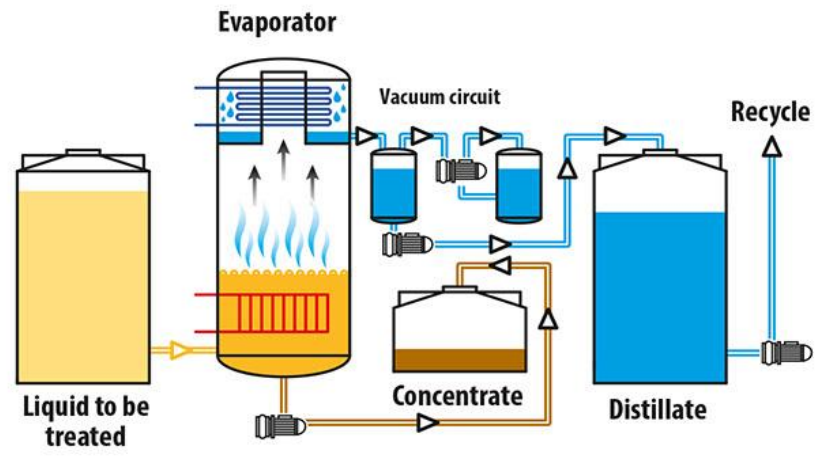


Figure 17: [Source: <http://www.pf10.com/evaporators-and-concentration-units.html>]

Chemical Oxygen Demand (COD) (Rozzi et al., 1996).

3.2.4.

(flocculation)

...

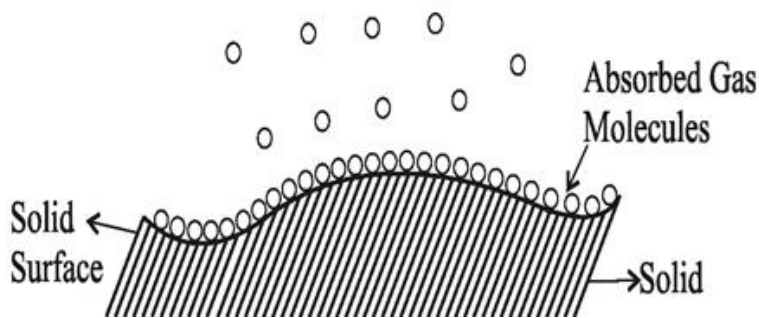
(Brenes et al., 1993; , 2010).

3.2.6.

(oxidation / reduction and detoxification) μ (H_2O_2) μ COD BOD₅ (, 2009).

3.2.7.

(adsorption)

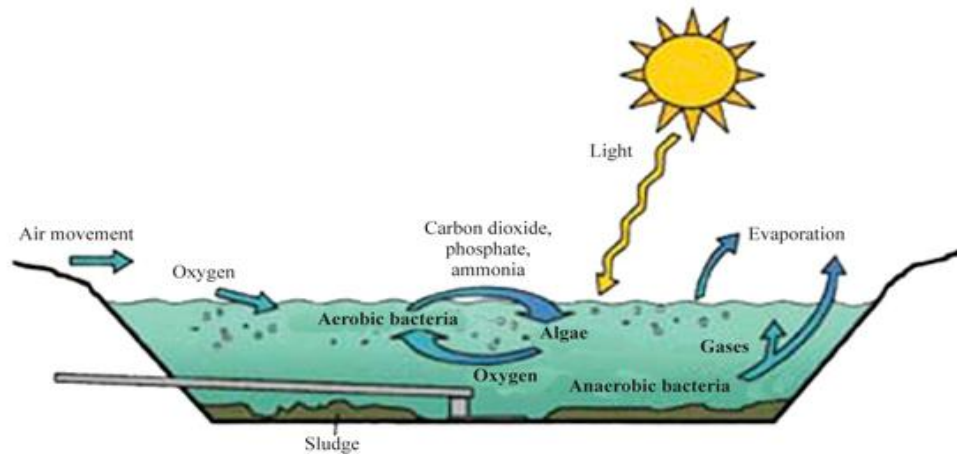


μ 19: [: <http://chemistry-desk.gr/2012/10/adsorption.html>]

3.3.

- (Lagoons)
- (Activated Sludge)
- (Anaerobic treatment)

3.3.1.



20: <http://extensionpublications.unl.edu/assets/html/>

...

1m³ 1m². (Rozzi et al., 1996).

(, 2010).

60 μ .

(Rozzi et al., 1996 ; , 2009).

3.3.2.

(activated sludge) (Drysdale et al., 1999). (2009) (biocenosis)

(2009 ; , 2011)

(NO³⁻). μ

0-0,5 ppm) (Drysdale et al., 1999).

(Drysdale et al., 1999).

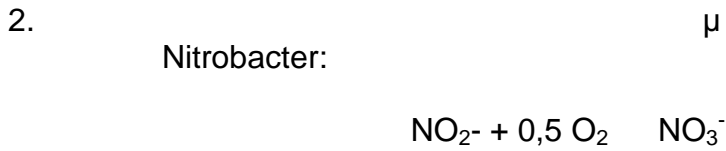
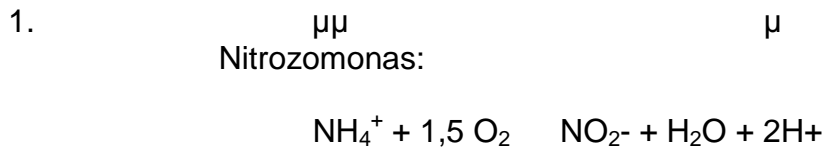
(nitrification)

(denitrification).

...

(Drysdale et al., 1999; Schlegel, 1992).

(Schlegel, 1992; Drysdale et al., 1999;):



(Schlegel, 1992 ; Drysdale et al., 1999;)



➤ (Lemmer et al., 1996). (2009)

(Ca(OH)₂, FeCl₃·O) (, 2009 ; , 2010).

(Lemmer et al., 1996).

➤ (Lacko et al., 1999):

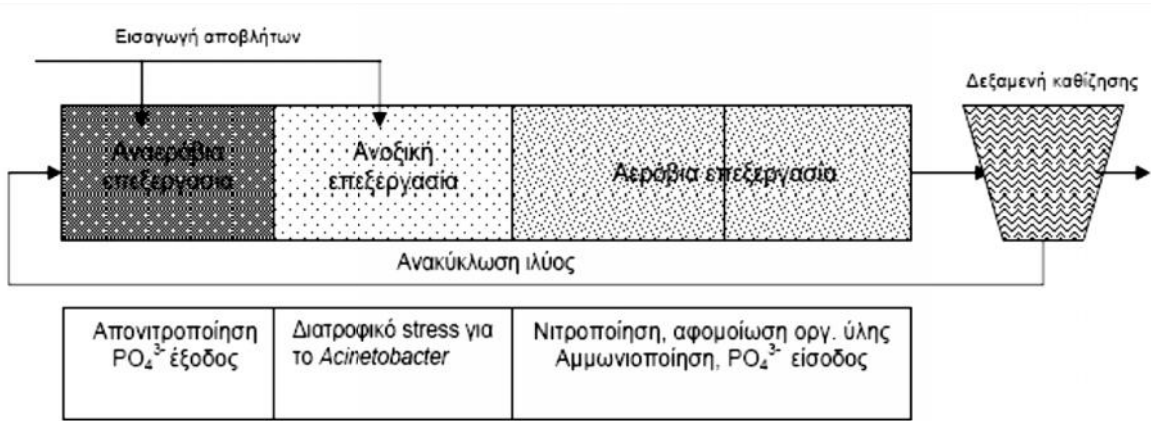
•

•

•

...

(Lacko et al., 1999).
 (Dabert et al, 2002).



μ 21: μ μ μ μ
 [: , 2009]

(MBRs), μ (microfiltration) (ultrafiltration).
 COD. μ BOD < 3000 mg/L
 mg/L (Brenes et al., 1993). BOD₅, μ 20.000 35.000
 CO₂
 () (, 2014 ;
 , 2009).

3.3.3.

(anaerobic treatment)

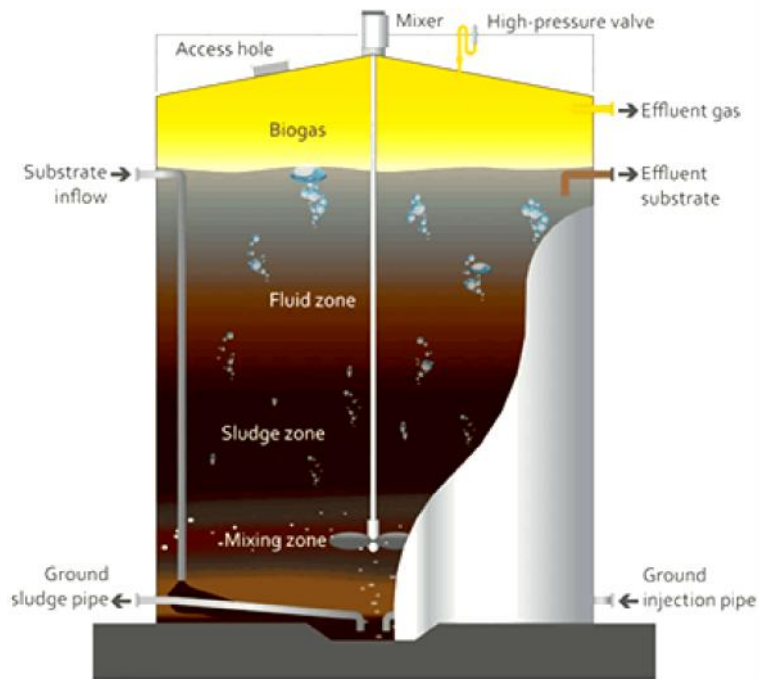


Figure 22: [Source: <http://www.mannvit.com/services/anaerobic-digestion/>]

The process involves the conversion of organic matter into biogas and sludge. The biogas is primarily composed of methane and carbon dioxide. The sludge is rich in organic matter and can be used as a fertilizer or as a substrate for further treatment. The process is highly efficient and can handle a wide range of organic waste. The digester is designed to maintain a constant temperature and pH, which are essential for the anaerobic process. The substrate is typically a mixture of organic waste and water. The digester is equipped with a mixer to ensure that the substrate is evenly distributed and that the biogas is collected. The effluent gas is collected and can be used for energy production. The effluent substrate is collected and can be used for various purposes. The ground sludge pipe is used to collect the sludge from the bottom of the digester. The ground injection pipe is used to inject water or other substances into the digester. The digester is a key component of an anaerobic treatment system and is used to treat a wide range of organic waste.

COD BOD₅ (7 g/L),

...

ΜΕΡΟΣ Β'

μ

4.

4.1.

(Daufin, 2001; Pizzichini & Russo, 2005; Fillaudeau et al., 2006; Mekonnen & Hoekstra, 2011). (Dhaouadi & Marrot, 2008).

(Gonçaves et al., 2009 ; Stoller & Bravi, 2010). (D'Annibale (2004)

30 μm^3 . (Paraskeva et al., 2007). 100-120 100
 (Di Giovacchino & Mascolo, 1988 ; Vlyssides et al., 1996 ; Gebreyohannes et al., 2015). (BOD₅) 45,5-68,7 g/lit, (COD) 85,7-158 g/lit (Di Giovacchino & Mascolo, 1988 ; Vlyssides et al., 1996).

(Roig et al., 2006 ; Gebreyohannes et al., 2015).

...



13: 0,10 gr
 TiO_2



14: 0,50 gr
 TiO_2



15: 0,80 gr TiO_2



16: μ μ .

...

12
 μ TiO₂ μ μ μ Ti₂-Degussa -25
 μ μ Degussa -25
 μ 50 m² / g. 80:20, μ μ 20 nm
12: μ μ

μ μ	μ	TiO ₂ (mg)	(ml)	μ /
				(mg/ml)
1	OMW1	100	200	0.5
2	OMW2	500	200	2.5
3	OMW3	1500	200	7.5
4	OMW4	3000	200	15

(17), μ μ μ μ 4 Watt. μ μ
 μ μ μ μ μ μ
 μ μ μ μ μ μ
 μ μ μ μ μ μ



17: μ μ μ
 μ μ

. . .

200 ml μ μ 300 . μ (SM) 5310 (μ μ (TOC)
 μ 0,9 mWatt / cm². μ PMA-2100 UV-Photometer
 μ , American Water Works Association) (Eaton, 2005 ;
Bekiari & Avramidis, 2014). μ μ
Shimadzu TOC (TOC-VCSH). μ μ

5.

TiO₂ is a semiconductor material that can be used for photocatalysis. When UV light is incident on TiO₂, it generates electron-hole pairs (e⁻ and h⁺). These pairs can react with water and oxygen to produce hydroxyl radicals (OH[•]) and superoxide anions (O₂^{•-}), which are highly reactive species that can degrade organic pollutants. (Mahlambi, Ngila & Mamba, 2015 ; Ibhaddon & Fitzpatrick, 2013).

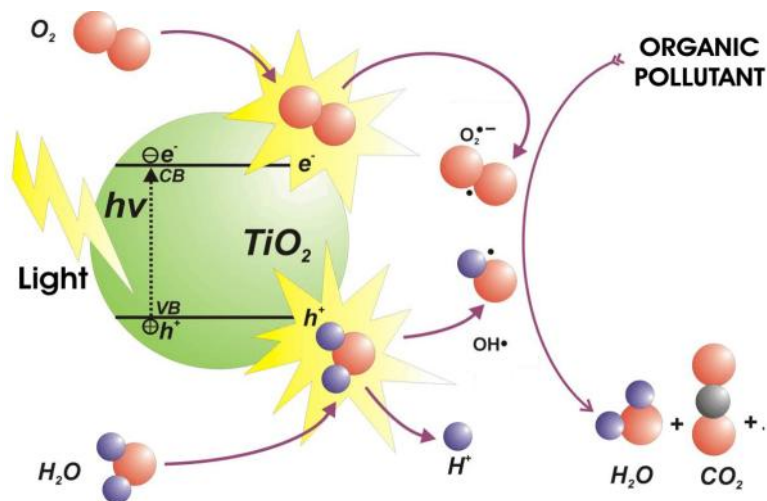


Figure 23: Photocatalytic mechanism of TiO₂. [Source: Ibhaddon & Fitzpatrick, 2013]

The photocatalytic mechanism of TiO₂ involves the generation of electron-hole pairs (e⁻ and h⁺) under UV light. The electron (e⁻) reacts with O₂ to form superoxide anions (O₂^{•-}), and the hole (h⁺) reacts with H₂O to form hydroxyl radicals (OH[•]) and protons (H⁺). These species then react with organic pollutants to produce H₂O and CO₂. The rate of photocatalysis is dependent on the intensity of light and the concentration of the pollutant. (Ibhaddon & Fitzpatrick, 2013 ; Mahlambi et al., 2015).

(+2,8 Volt),

...

H_2O_2 (+1,78 Volt) O_3 (+2,07 Volt), $HOCl$ (+1,49 Volt)
 (Ibhadon & Fitzpatrick, 2013 ; Mahlambi et al., 2015).

TiO_2 (UV) (380 nm, 400 nm).
 UV (Xiaoyong Wu et al., 2013).

οργς (TOC) :

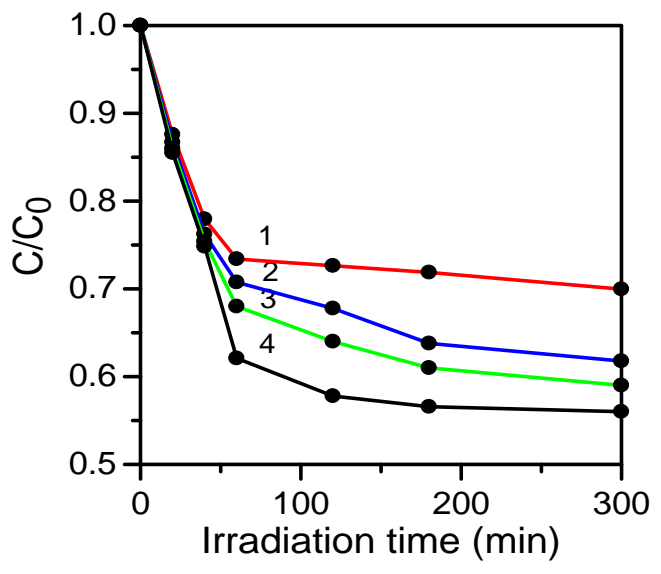
$$R = \frac{C_0 - C}{C_0} \quad [1]$$

C_0 : TOC μ
 C : TOC μ

στη συν (E%) :

$$E\% = \left(\frac{C_0 - C}{C_0} \right) * 100\% \quad [2]$$

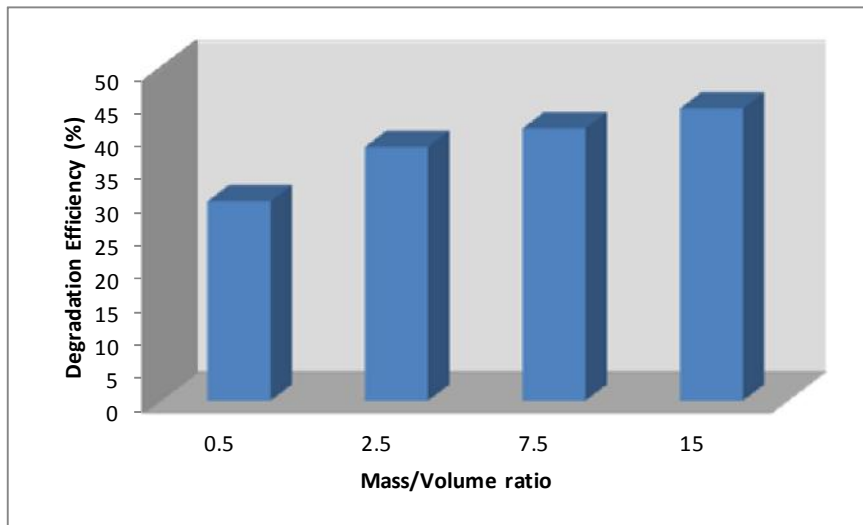
TiO_2 , TOC, μ 24).



μ 24: TOC μ

...

TOC (3).



25: 300

TOC / 44%
30%, (UV) 300 (22, 25).

. . ., 2003, , ,
μ ., 1993, μ , μ μ ,
μ μ -, μ .
. 1994, - μ μ -
μ μ . μ μ ,
μ ., 2011, μ , . .
, ,
, 2014, μ μ μ μ μ
, μ μ , μ μ
, . μ
. , 2009, (pabr) μ ,
, μ μ , μ μ ,
. , 1983, ,
μ ., 2008, ,
μ , , 1986, - - ,
μ
. , 1994, . ,
. , 2010, μ μ
, μ , μ
,

. . .

., 2006,
μ , , μ ,
., 1999,
: μ ,
& , 14, p.52-59.
., 2000, μ μ , μ ,
μ ,
., 1994, μ : - ,
2 .
., 1990, μ , μ ,
., 1999, μ , μ ,

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. . .

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ΠΑΡΑΡΤΗΜΑ Α

OLIVE MILL WASTEWATERS TOTAL ORGANIC
CARBON DEGRADATION USING TiO_2
NANOPARTICLES

Olive Mill Wastewaters Total Organic Carbon Degradation using TiO₂ Nanoparticles

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Abstract

Olive mill wastewaters (OMW) constitute an important pollution factor for the olive oil-producing regions but also a significant problem to be solved for the agricultural industry. The main reason is the large wastes amounts produced in relatively small time interval, which should be processed with safety for the environment.

Because of the high organic content, it is imperative to use Advanced Oxidation Technologies in order to reduce the organic load of these wastewaters.

We used TiO₂ nanopowder (Degussa P-25) as a low cost, low toxicity and effective photocatalyst for the degradation of the organic load of the wastewaters in an olive oil production facility at the prefecture of Ilea Western Greece. The treatment of OMW showed that the Total Organic Carbon (TOC) can be reduced over 44% after 300 minutes under UV irradiation. In addition we studied the effect of the TiO₂ mass in respect to the total volume ratio of OMW as a critical factor for the effective degradation of the TOC. The sample with a mass/volume = 15 mg/ml shows the highest photocatalytic decomposition efficiency (44%) in comparison to the mass/volume = 0.5 mg/ml which shows a 30% efficiency after 300 minutes of UV exposure.

This easy-going treatment technology, aim to transform resistant organic molecules into others which could be further biodegraded in the natural environment.

Keywords: Olive Mills Wastewater, Photocatalysis, TiO₂ nanopowder.

1. Introduction

Olive mill wastewaters (OMW) are of the major environmental problems in Mediterranean countries where large quantities of olive oil are produced (1-5). The acidic pH values, the high organic load and the large quantities are the major characteristics that constitute olive mills wastewaters, hazardous wastes (6).

The OMW produced during the treatment of the olives, in different stages, in traditional, and/or two-three phase centrifugal systems (7-8). An estimation of the OMW produced under these facilities in the Mediterranean is about 30 million m³ (9). However, the composition of the OMW exhibits variations in regard to the methodology of the oil extraction, the age of the trees, the soil characteristics and other environmental factors (10).

In general, for each 100 kilos of olive-crop, 100-120 kilos of humid wastes are produced.

The substantial high organic load constitutes by substances like sugars, organic acids, amino-acids, proteins, fats and polyphenols. Wastes contain high concentration of polyphenols, which may cause the appearance of bio-toxic phenomena in the natural environment.

A typical composition of the OMW is organic matter of about 4-16 %wt, minerals 1-2 %wt and water 83-92 %wt (11-13). Total phenols are between 10.6-17.2 g/l, BOD₅ 45.5-68.7 g/l and COD 85.7-158 g/l (11-12).

The illegal discharge of the OMW without pretreatment is of a major concern in oils produced countries in the Mediterranean region and causes a lot of environmental problems such as eutrophication, toxic phenomena to the aquatic fauna, phytotoxicity, aesthetic degradation, as well as socioeconomic impacts in a regional scale (13-15). During the past, much research has been conducted in order to achieve an efficient degradation method of the high organic load of the OMW.

These methodologies used various techniques such as the aerobic-anaerobic treatment (16-18), enzymatic catalysis (19), composting (20), membrane ultra-filtration (UF) combined with centrifugation (21) or UV/H₂O₂ oxidation (22), advance oxidation using O₃/UV (23), photocatalysis (24), Fenton and electrochemical oxidation (25, 26).

However, a major disadvantage of the above mention techniques is the up-scaling application, and their use in real conditions olive oil production units.

Table 1. Mass to volume ratio of the samples analyzed

Sample Number	Sample ID	Mass TiO ₂ (mg)	Volume Wastewater (ml)	Mass/Volume (mg/ml)
1	OMW1	100	200	0.5
2	OMW2	500	200	2.5
3	OMW3	1500	200	7.5
4	OMW4	3000	200	15

Taking into consideration the limitations of the use of advance oxidation technologies for OMW in large scale applications, we propose a simple methodology based on the photocatalytic degradation of the total organic load using TiO₂ as the photocatalyst for OMW effective treatment.

2. Methodology, Results and Discussion

2.1. Materials and Methodology

Samples were collected from a two phase centrifugal system, oil production facility, located in Pyrgos municipality, Ileia prefecture Western Greece. After collection, the samples were deep frozen until their treatment to the laboratory. The raw sample was processed to a 9 mesh, 2 mm pore size strainer (Fig. 1). The wastewater was then centrifuged two times (10 minutes each) using a speed of 3000 r/min. Aliquots of the sample were used for the photocatalysis.



Figure 1. Photographs showing the treatment of the samples analyzed.

In Table 1 we present the samples identification and the mass of TiO₂ nanopowder and volume of sample used. Commercially available TiO₂ -Degussa P-25 was used as the photocatalyst. The Degussa P-25 characteristics are anatase to rutile ratio 80:20, particle size 20 nm and particle specific area 50 m²/g.

A cylindrical reactor was used in all experiments (Fig. 2). Four black light fluorescent tubes of 4 W each nominal power, were placed around the reactor. The whole construction was covered with a cylindrical aluminum reflector. Cooling was achieved

by air flow from below the reactor using a ventilator. Continuously stirring of the samples in the reactor was achieved with a magnetic stirrer.



Figure 2. Photographs showing the photocatalytic system used.

The intensity of radiation was measured with a Solar Light PMA-2100 UV-Photometer and found equal with 0.9 mW/cm². The reactor was filled with 200 ml of the sample and the irradiation applied for a total of 300 minutes. TOC analysis was performed using the Combustion-Infrared method, Standard Method (SM) 5310B (Standard Methods for the Examination of Water and Waste Water, American Water Works Association) (27, 28). All analyses were carried out using a Shimadzu TOC analyzer (TOC-VCSH).

2.2. Results and Discussion

The use of TiO₂ powder as the photocatalyst was decided because titanium dioxide is a well-known low cost nontoxic photocatalyst for a variety of pollutants. Also it can be used with success in many photocatalytic cycles and can be easily handled by an unskilled worker in an oil production unit. The high efficiency of TiO₂ to produce hydroxyl radicals in solution under UV light illumination, the high stability in water and the nontoxic response are major characteristics. In addition TiO₂ is a cost efficient and very effective photocatalyst for the decomposition of organic substances in water and air (Fig. 3; 29, 30).

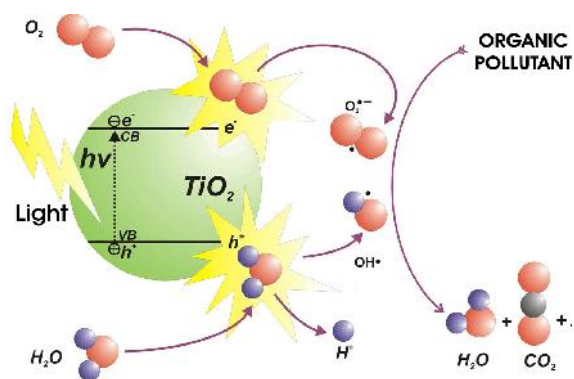


Figure 3. Schematic presentation of $\cdot\text{OH}$ generation through TiO_2 photoexcitation and organic pollutants degradation (adopted from 30).

Hydroxyl radicals ($\cdot\text{OH}$) can be generated from water using TiO_2 as a photocatalyst. The $\cdot\text{OH}$ compounds are very reactive chemicals with a redox potential of +2.8 V (vs. Nernst Hydrogen Electrode) and they can react with the organic pollutants with a $10^7\text{-}10^{10} \text{ M}^{-1} \text{ s}^{-1}$ constant reaction rate. The end product of the reaction of the $\cdot\text{OH}$ radicals with the organic compounds are CO_2 , H_2O and inorganic salts (29, 30). Because of hydroxyl radicals' high redox potential (+2.8 V), they are more effective for the decomposition of organic pollutants than other oxidants like O_3 (+2.07 V), HOCl (+1.49 V) and H_2O_2 (+1.78 V) used for water purification and disinfection (29, 30).

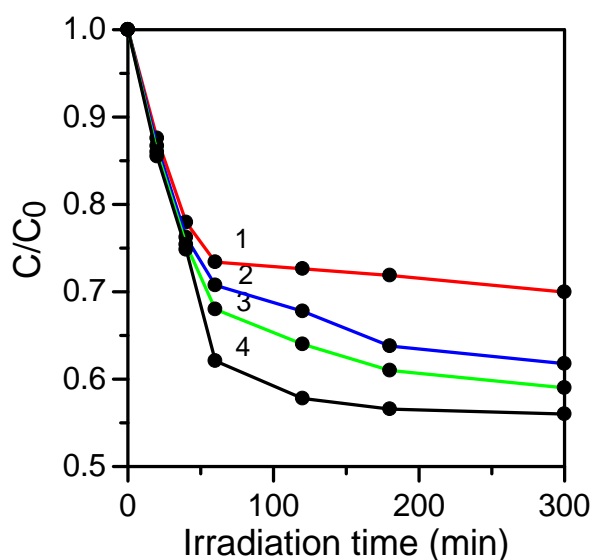


Figure 4. Decomposition rate of TOC in relation to the irradiation time.

The effectiveness of semiconductors for photoactivation is a function of the energy required for the excitation of their crystals. For TiO_2 in the anatase form this energy must be higher than the $E_g=3.2 \text{ eV}$ and for the rutile higher than $E_g=3.0 \text{ eV}$. As a conclusion TiO_2 nanocrystals excitation requires

energy in the near UV region (radiation with 380 nm for anatase and 400 nm for rutile). This limitation is a major disadvantage of the use of TiO_2 as a photocatalyst because only 5% of the solar light radiation is in the UV region (31).

In order to estimate the photo-degradation rate (R) of the TOC we employ the following equation:

$$R = (C_0 - C) / C_0$$

Where C_0 is the initial concentration of the pollutant measured in solution and C is the final concentration after irradiation with UV light. Then we can calculate the degradation efficiency (E%) as:

$$E\% = [(C_0 - C) / C_0] \times 100\%$$

Our experiments revealed that the photocatalytic efficiency of the TiO_2 used for the degradation of the TOC in the samples analyzed is related to the time of irradiation and the mass of the material used. In particular after more than 150 minutes of illumination all the samples showed no difference in the photo-degradation rate of TOC (Fig. 4).

The mass of the catalyst to the total volume of the sample, indicates that the higher mass/volume ratio causes the most efficient degradation of the TOC among all measured samples (Fig. 5).

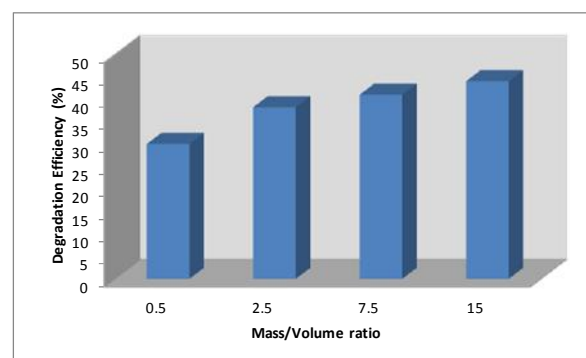


Figure 5. Degradation efficiency of the catalyst after 300 min of irradiation with respect to the mass/volume ratio.

High TOC decomposition efficiency 44% is apparent when we use the higher mass/volume ratio while the lowest mass/volume ratio shows a degradation efficiency of 30% after 300 min of illumination (Fig. 5). Higher mass/volume ratio could be used and repetition measurements could be helpful.

3. Conclusions

The data presented here show that TiO_2 can be efficiently used for the photocatalytic treatment of Olive Mills Wastewaters. The degradation efficiency of the photocatalyst used for the decomposition of the

OLIVE MILL WASTEWATERS ORGANIC POLLUTANTS DEGRADATION USING TiO₂ NANOPARTICLES

Total Organic Carbon of the samples is accelerated when high mass TiO₂/volume sample ratios have been used. After 150 minutes of UV irradiation there is no change of the decomposition rate of TOC.

Degussa P-25 can be used as a low cost, efficient and environmental friendly material for the treatment of the olive mills wastewaters generated in an olive oil production facility. The procedure we propose for an easy going process is illustrated in figure 6. To this perspective, the OMW treatment includes physical sedimentation in the first concrete tank in order for the heavy solid particles to precipitate and a second step with stirring, oxygenation and photocatalytic treatment with Degussa P-25 under UV illumination for 150 minutes.

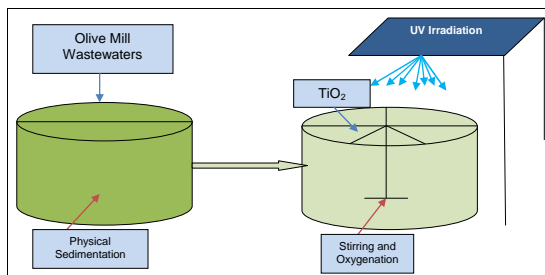


Figure 6. Schematic illustration of a cost effective, efficient and easy going procedure for OMW treatment.

Oxygenation can further facilitate the TOC photocatalytic degradation because of the major reactive oxygen species (ROS) produced (32).

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Dr. Dionisios Panagiotaras

Από: cest@gnest.org
Αποστολή: Τρίτη, 25 Απριλίου 2017 4:13 μμ
Προς: sakpanag@teiwest.gr
Θέμα: [CEST] New decision has been made

Dear Dr. Dionisios Panagiotaras

On behalf of the Scientific Committee of the 15th International Conference on Environmental Science and Technology (CEST2017), we are pleased to inform you that your paper

Ref No: cest2017_00589

Title: Olive Mill Wastewaters Total Organic Carbon Degradation using TiO₂ Nanoparticles

has been accepted for oral presentation. The evaluation was based on the feedback that we received from the reviewers taking into consideration the quality and the relevance of your paper.

Thanking you for your contribution, we are looking forward to seeing you in Rhodes, Greece.

Yours sincerely,

Dr Demetris F. Lekkas

On behalf of the Scientific and Programme Committee

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μ (mg/L)	40	35	50	35	500	80
(mg/L)	5	25	5	25	40	100
(mg/L)	0,5	0,1	0,5	0,1	5	10
(mg/L)					25	
(mg/L)					10	
(C)		35		35	35	45
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75/442/		96/350,		75/442/ 15	75/442/ : 49541/1424/86	15	1980	μ	μ
91/156/	μ	μ	75/442	μμ 91/156:	69728/824/96	μ	μ	μ	91/692/ C
					114218/1997		μμ	μ	96/350/ μ
					18186/271/1988			μ μ	
76/464/EEC	μ	μ	1976	76/464/ 4	4859/726/2001			μ	
				μ	μ μ μ	4	1976	μ	μ
					76/464/ . . 2/1.2.2001		μ	μ	μ
					μ	μ	μ	μ	μ
				4	1976	μ	μ	μ	μ
80/68/	μ	μ	1979	80/68/ 17	26857/553/1988			μ	μ
				μ	μ			μ	μ
86/280/					86/280/ : 55648/2210/91			μ	μ
90/415/	88/347/ μ			86/280/ μ	μ		88/347/ μ		
				90/415/ μ	90/415/ 90461/2193/94	μ	12	μ	55648/2210/91
				μ	255/94	μ	12	μ	55648/2210/91
				76/464/ μ	88/347/ μ		μ	μ	6 μ. 144/2.11.1987
259/93/				μ ()	μ			μ	μ
				259/93 1	1993			μ	μ
				μ	μ			μ	μ
				μ	μ			μ	μ

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1. ., 2006, , , μ μ μ ,
 2. μ ., 2011, μ , . . .
 3. :
- <http://www.elinyae.gr/el/>